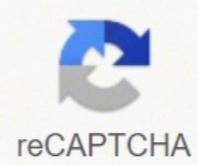


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Mole Calculation Practice Worksheet

water = $H_2O \rightarrow 2(1.0) + 16.00 = 18.02\text{ g}$

1) How many moles are in 25 grams of water?
 $25\text{g} H_2O \times \frac{1\text{mol}}{18.02\text{g}} = 1.39\text{ moles } H_2O$

1a) How many moles are in 10g of NaBr?

1b) How many moles are in 200g of PbSO₄?

COVALENT BONDING						
Draw Lewis structures for each molecule below. Circle all compounds that obey the octet rule and put a star next to all compounds that violate the octet rule. Indicate the bonding regions and unshared pairs for the central atom (or least electronegative excluding hydrogen) and tell how many sigma and pi bonds each compound has.						
molecular geometry	Structure	Bonding Regions	Unshared Pairs	Sigma Bonds	Pi Bonds	hybridization
octahedral		6	0	6	0	sp ³ d ²
linear		2	0	2	2	sp
	$\text{NO}_2 : \ddot{\text{O}}-\ddot{\text{N}}=\ddot{\text{O}} \leftrightarrow \ddot{\text{O}}=\ddot{\text{N}}-\ddot{\text{O}}$	2	1/2	2	1	
bent	$[\ddot{\text{O}}-\ddot{\text{N}}=\ddot{\text{O}}] \leftrightarrow [\ddot{\text{O}}=\ddot{\text{N}}-\ddot{\text{O}}]$	2	1	2	1	sp ²
trigonal planar		3	0	3	0	sp ²
trigonal pyramidal		3	1	3	0	sp ³
seesaw	$[\ddot{\text{F}}-\ddot{\text{As}}-\ddot{\text{F}}]^-$	4	1	4	0	sp ³ d ²
tetrahedral	$[\ddot{\text{O}}-\ddot{\text{S}}-\ddot{\text{O}}]^2$	4	0	4	0	sp ³

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Masses and Moles

Look at this example of how to calculate the mass from the number of moles:

$$\text{Mass of substance} = \text{number of moles} \times \text{mass of 1 mole}$$

Which is the same as saying...
 $\text{Mass} = \text{number of moles} \times \text{Relative Formula Mass}$

$$m = n \times M_r$$

**Example** What is the mass of 2 moles of carbon?

m	=	$n \times M_r$
	=	2×12
	=	24g

Use the Periodic Table in the front of the book to answer these questions:

Q1 What is the mass of...?

- a) 1 mole of carbon atoms
- b) 2 moles of sodium atoms
- c) 2 moles of aluminium atoms
- d) 3 moles of lithium atoms
- e) 10 moles of iodine atoms
- f) 1 mole of oxygen molecules
- g) 2 moles of chlorine molecules
- h) 1 mole of nitrogen molecules
- i) 0.5 moles of nitrogen molecules
- j) 0.1 moles of carbon molecules

Q2 What is the mass of...?

- a) 1 mole of carbon dioxide (CO_2)
- b) 5 moles of carbon dioxide (CO_2)
- c) 3 moles of water (H_2O)
- d) 2 moles of sulphur trioxide (SO_3)
- e) 3 moles of nitric acid (HNO_3)
- f) 10 moles of hydrochloric acid (HCl)
- g) 100 moles of sulphuric acid (H_2SO_4)
- h) 50 moles of sodium hydroxide (NaOH)
- i) 30 moles of calcium oxide (CaO)
- j) 25 moles of sodium chloride (NaCl)

Q3 What is the mass of...?

- a) 0.1 moles of copper oxide (CuO)
- b) 0.1 moles of calcium carbonate (CaCO_3)
- c) 0.01 moles of sulphuric acid (H_2SO_4)
- d) 0.2 moles of calcium chloride (CaCl_2)
- e) 0.25 moles of sodium carbonate (Na_2CO_3)
- f) 0.05 moles of sodium hydroxide (NaOH)
- g) 0.58 moles of aluminium oxide (Al_2O_3)
- h) 0.1 moles of propanoic acid ($\text{C}_3\text{H}_5\text{COOH}$)
- i) 0.5 moles of potassium manganate(VII) (KMnO_4)
- j) 0.25 moles of Hydrated Magnesium Sulphate ($(\text{MgSO}_4 \cdot 7\text{H}_2\text{O})$)

Section Three — Equations

THE MOLE AND
AVOGADRO'S NUMBER

Name _____

One mole of a substance contains Avogadro's Number (6.02×10^{23}) of molecules.

How many molecules are in the quantities below?

1. 2.0 moles
2. 1.5 moles
3. 0.75 mole
4. 15 moles
5. 0.35 mole

How many moles are in the number of molecules below?

1. 6.02×10^{23}
2. 1.204×10^{24}
3. 1.5×10^{23}
4. 3.4×10^{24}
5. 7.5×10^{23}

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