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Mole Calculation Practice Worksheet

water = $H_2O \rightarrow 2(1.01) + 16.00 = 18.02g$
mole

1) How many moles are in 25 grams of water?

$25g H_2O \times \frac{1mol}{18.02g} = 1.39moles H_2O$

1a) How many moles are in 10g of NaBr?

1b) How many moles are in 200g of $PbSO_4$?

COVALENT BONDING

Draw Lewis structures for each molecule below. Circle all compounds that obey the octet rule and put a star next to all compounds that violate the octet rule. Indicate the bonding regions and unshared pairs for the **most electronegative** (or least electronegative excluding hydrogen) and tell how many sigma and pi bonds each compound has.

molecular geometry	Molecule	Structure	Bonding Regions	Unshared Pairs	Sigma Bonds	Pi Bonds	hybridization
octahedral	SCl_6		6	0	6	0	sp^3d^2
linear	N_2		2	0	2	2	sp
	NO_2		2	1/2	2	1	
bent	NO_2^+		2	1	2	1	sp^2
trigonal planar	BF_3		3	0	3	0	sp^2
trigonal pyramidal	PH_3		3	1	3	0	sp^3
seesaw	AsF_5		4	1	4	0	sp^3d
tetrahedral	SO_4^{2-}		4	0	4	0	sp^3

Masses and Moles

Look at this example of how to calculate the mass from the number of moles:

Mass of substance = number of moles x mass of 1 mole

Which is the same as saying....

Mass = number of moles x Relative Formula Mass

$m = n \times M_r$

$\frac{m}{n \times M_r}$

Example

What is the mass of 2 moles of carbon?

$m = n \times M_r$
 $= 2 \times 12$
 $= 24g$

Use the Periodic Table in the front of the book to answer these questions:

Q1 What is the mass of...?

- a) 1 mole of carbon atoms
- b) 2 moles of sodium atoms
- c) 2 moles of aluminium atoms
- d) 3 moles of lithium atoms
- e) 10 moles of iodine atoms
- f) 1 mole of oxygen molecules
- g) 2 moles of chlorine molecules
- h) 1 mole of nitrogen molecules
- i) 0.5 moles of nitrogen molecules
- j) 0.1 moles of carbon molecules

Q2 What is the mass of...?

- a) 1 mole of carbon dioxide (CO_2)
- b) 5 moles of carbon dioxide (CO_2)
- c) 3 moles of water (H_2O)
- d) 2 moles of sulphur trioxide (SO_3)
- e) 3 moles of nitric acid (HNO_3)
- f) 10 moles of hydrochloric acid (HCl)
- g) 100 moles of sulphuric acid (H_2SO_4)
- h) 50 moles of sodium hydroxide ($NaOH$)
- i) 30 moles of calcium oxide (CaO)
- j) 25 moles of sodium chloride ($NaCl$)

Q3 What is the mass of...?

- a) 0.1 moles of copper oxide (CuO)
- b) 0.1 moles of calcium carbonate ($CaCO_3$)
- c) 0.01 moles of sulphuric acid (H_2SO_4)
- d) 0.2 moles of calcium chloride ($CaCl_2$)
- e) 0.25 moles of sodium carbonate (Na_2CO_3)
- f) 0.05 moles of sodium hydroxide ($NaOH$)
- g) 0.58 moles of aluminium oxide (Al_2O_3)
- h) 0.1 moles of propanoic acid (C_2H_5COOH)
- i) 0.5 moles of potassium manganate(VII) ($KMnO_4$)
- j) 0.25 moles of Hydrated Magnesium Sulphate ($MgSO_4 \cdot 7H_2O$)

Section Three — Equations

THE MOLE AND AVOGADRO'S NUMBER

Name _____

One mole of a substance contains Avogadro's Number (6.02×10^{23}) of molecules.

How many molecules are in the quantities below?

1. 2.0 moles
2. 1.5 moles
3. 0.75 mole
4. 15 moles
5. 0.35 mole

How many moles are in the number of molecules below?

1. 6.02×10^{23}
2. 1.204×10^{24}
3. 1.5×10^{24}
4. 3.4×10^{24}
5. 7.5×10^{24}

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